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grandtheftauto5activationcodeBile acid binding in the ion pair of 1,1'-diethyl-3,3,3',3'-tetramethyl-2,2'-bipyridinium bromide-water and protonation of the ion pair of tauro-taurocholic acid-acetate. The bile acids bind to the weak polyanion 1,1'-diethyl-3,3,3',3'-tetramethyl-2,2'-bipyridinium bromide in the water/bile acid ion pair. The pH-dependent equilibrium behavior of the bile acid in aqueous 1,1'-diethyl-3,3,3',3'-tetramethyl-2,2'-bipyridinium bromide solution has been investigated by spectrophotometry at 303 K, using optical absorption and circular dichroism spectroscopies. The weak polyacid 1,1'-diethyl-3,3,3',3'-tetramethyl-2,2'-bipyridinium bromide acts as a complexing agent for bile acids, such that the parent molecule aqueous solution is practically neutral, whereas the titration of the bile acid makes the polyacid protonated, and the equilibrium shifts to the bile acid containing ion pair. The presence of an organic solvent in the bile acid solution plays a key role in the process of protonation; a smaller pK of the bile acid in ethanol than in H₂O can explain this behavior. The stability constant (log K) obtained for each bile acid is very similar for both acetic and acetate acids, and the lack of differences suggests that the equilibria investigated are essentially stoichiometric. The enthalpy of protonation of the bile acid ($\Delta H(p)$ degrees) and the log K for the ion pair is strongly dependent on the bile acid structure. The values of $\Delta H(p)$ degrees increase with the number of hydroxyl groups in the tauroside chain and decrease with the number of carbon atoms in the side chain. An almost linear dependence of log K for the ion pair on the Hammett rho (Hammett) constant is obtained, revealing the rho value of the bile acid acid as about -0.5. It's clear

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